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Calculations:
To find out the percentage purity of $KMnO_4$

Working formula:
$$\frac{M_1 V_1}{M_2 V_2} = \frac{1}{5}$$
, where, M_1 = Molarity of $KMnO_4$ solution
 M_2 = Molarity of Mohr's salt soln.
 V_1 = Volume of $KMnO_4$ solution
 V_2 = Volume of Mohr's salt solution

Given: Strength in grams per litre of $KMnO_4$ solution = 2.75
Strength in g/l of Mohr's salt solution = 20.5

$M_2 = \frac{\text{Strength in g/l of Mohr's salt solution}}{\text{Mol. wt. of Mohr's salt}}$

$= \frac{20.5}{392} = 0.05229$

$\frac{M_1 V_1}{M_2 V_2} = \frac{1}{5}$

$\Rightarrow M_1 = \frac{M_2 V_2}{5 V_1} = \frac{0.05229 \times 25}{5 \times 16.8} = 0.01556$

$\therefore M_1 = 0.01556$

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